

solutions and colligative properties pdf

Solutions and Colligative Properties Introduction: Solutions: Mixture of two or more components. Depending on size of components, mixtures are classified into 3 types. a) Coarse mixture: Definition: The mixture which contains components having relatively bigger size is called as coarse mixture. E.g. Mixture of salt and sugar.

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Colligative Properties Vapour pressure Boiling point Freezing point Osmotic pressure. Learning objectives Describe meaning of colligative property Use Raoult's law to determine vapor pressure of solutions ...

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Colligative Properties- Page 1 Lecture 4: Colligative Properties By definition a colligative property is a solution property (a property of mixtures) for which it is the amount of solute dissolved in the solvent matters but the kind of solute does not matter.

Colligative Properties- Page 1 Lecture 4: Colligative

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Colligative properties of solutions Colligative properties are those which depend on the concentration of the solution. These properties do not depend on the nature or type of solute but only on the amount of solute present in the solution. This is the reason why we will only consider non- ...

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11. Osmotic pressure The osmotic pressure for an ideal dilute solution is approximated by the equation: $\pi = MRT$ where π is the osmotic pressure (in atm), M is the molarity of the solute, R is the gas constant (0.0821 L-atm/mol-K) and T is the absolute temperature.

SOLUTIONS AND THEIR COLLIGATIVE PROPERTIES

Solutions 6.6 Colligative Properties of Solutions Worksheet 1) A student dissolves 20.0 g of glucose, $C_6H_{12}O_6$, into 511 mL of water at 25 °C. The vapor pressure of pure water at 25°C is 3.13×10^{-2} atm. Find the vapor pressure of the solution.

Solutions 6.6 Colligative Properties of Solutions Worksheet

Colligative Properties of Solutions Liquids solutions experience the following four colligative properties: Vapor Pressure Reduction, Boiling Point Elevation,

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